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Electrochemistry: Crash Course Chemistry #36
Electrochemistry

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Reaction | IIT JEE Main \u0026 Advanced |
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Nernst Equation + Example (Concentrations)
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AP Chemistry-Electrochemistry. Multiple Choice. Identify the choice that best completes the statement or answers the question. ____ 1. The half-reaction that occurs at the cathode during the electrolysis of molten sodium bromide is _____. a. $+ 2e^- 2Br^- \rightarrow Br_2$ b. $+ 2e^- Br_2 \rightarrow 2Br^-$ c. $+ e^- Na^+ \rightarrow Na$ d. $Na \rightarrow Na^+ + e^-$ e. $\rightarrow 2H_2O + 2e^- 2OH^- + H_2$ ____

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2.

AP Chemistry-Electrochemistry - Quia
AP Chemistry: Electrochemistry Multiple
Choice Answers 14. Questions 14-17 The
spontaneous reaction that occurs when the
cell in the picture operates is as follows:
 $2\text{Ag}^+ + \text{Cd (s)} \rightarrow 2 \text{Ag (s)} + \text{Cd}^{2+}$ (A) Voltage
increases. (B) Voltage decreases but remains
> zero.

AP Chemistry: Electrochemistry Multiple

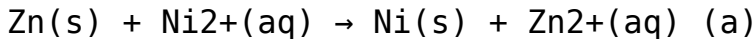
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Identify M and M^{2+} in the diagram and specify the initial concentration for M^{2+} in solution. Electrons flow from the anode to the cathode in a voltaic electrochemical cell. The anode is where oxidation occurs, and in the reaction above, $Zn(s)$ is oxidized.

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1996 Free Response Questions 7) $Sr(s) + Mg^{2+}$
 $\rightleftharpoons Sr + Mg(s)$ Consider the reaction

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represented above that occurs at 25°C . All reactants and products are in their standard states.

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AP REVIEW QUESTIONS – Electrochemistry -
Answers Answer: (a) tin electrode is the cathode; cathode is the site of reduction (gain in electrons) and will convert metal ions into a metal. (b) (see diagram) (c) red: $\text{Sn}^{2+} (\text{aq}) + 2 \text{e}^- \rightarrow \text{Sn} (\text{s})$ $E^\circ = -0.14 \text{ V}$ oxid: $\text{X} (\text{s}) - 3 \text{e}^- \rightarrow \text{X}^{3+} (\text{aq})$ $E^\circ = +0.74 \text{ V}$ $E^\circ \text{ cell} = +0.60 \text{ V}$ red: $\text{X}^{3+} (\text{aq}) + 3 \text{e}^- \rightarrow \text{X}$

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Response Questions 7) $\text{Sr}(s) + \text{Mg}^{2+} \rightleftharpoons \text{Sr} + \text{Mg}(s)$ Consider the reaction represented above that occurs at 25°C . All reactants and products are in their standard states. The value of the equilibrium constant, K_{eq} , for the reaction is 4.2×10^{17} at 25°C .

A.P. Chemistry Practice Test - Ch. 17:
Electrochemistry A ...

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Electrochemistry. Redox reaction from dissolving zinc in copper sulfate.

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Electrochemistry Involves TWO MAIN TYPES Of
Electrochemical Cells : 1. Galvanic (voltaic)
cells – which are thermodynamically favorable
chemical reactions (battery) 2. Electrolytic
cells – which are thermodynamically
unfavorable and require external e⁻ source (a
direct current or DC power source)

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AP Chemistry Review Questions -

Electrochemistry. For the galvanic cell

described below, the correct line notation

is: $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$ ($E^\ominus = 1.36\text{v}$) $\text{Cu} + \text{e}^- \rightarrow \text{Cu}$ ($E^\ominus = 0.52\text{v}$) $\text{Cu (s)} | \text{Cu}^+ (\text{aq}) || \text{Cl}_2 (\text{g}) | 2\text{Cl}^- (\text{aq}) | \text{Pt (s)}$ $\text{Pt (s)} | \text{Cu (s)} | \text{Cu}^+ (\text{aq}) || \text{Cl}_2 (\text{g}) | 2\text{Cl}^- (\text{aq}) | \text{Pt (s)}$

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Answer the following questions regarding the electrochemical cell shown above. (a) Write the balanced net-ionic equation for the spontaneous reaction that occurs as the cell operates, and determine the cell voltage. (b) In which direction do anions flow in the salt bridge as the cell operates? Justify your answer. (c) If 10.0 mL of 3.0-molar AgNO_3

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