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true 2. true 3. false 4.

true 5. true 6. 2, 2, 64.10

7. 3, 3, 96.00 8. 2, 2,

88.02 9. 4, 4, 72.08 10.

methanol and oxygen gas 11.

carbon dioxide and water 12.

160.10 g 13. 160.10 g 14.

They are equal. 15. A mole

ratio is a ratio between the
numbers of moles

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368 Chapter 11 •

Stoichiometry Section

11.11.1 Objectives Describe

the types of relationships

indicated by a balanced

chemical equation. State the

mole ratios from a balanced

chemical equation. Review

Vocabulary reactant: the

starting substance in a

chemical reaction New

Vocabulary stoichiometry

mole ratio Defining

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stoichiometry. mole ratio.

limiting reactant. excess

reactant. the study of

quantitative relationships

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between the amounts of... in a balanced equation, the ratio between the number of moles... a reactant that is totally consumed during a chemical reaction...

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11 209
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CHAPTER 11 SOLUTIONS MANUAL
Section 11.1 Defining
Stoichiometry pages 368–372
Practice Problems pages
371–372 1. Interpret the
following balanced chemical
equations in terms of
particles, moles, and mass.

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Show that the law of
conservation of mass is

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Class Stoichiometry Section
11.1 What is stoichiometry?
In your textbook, read about

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stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction

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In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each

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(recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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to express the stoichiometry
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the numbers

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sinned willfully, trampled
under foot the Son of God,
regarded as unclean the
blood ...

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