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Answer Key Chapter 12: Stoichiometry Mole Ratios Questions 1. Aluminum reacts with oxygen to produce aluminum oxide as follows:  $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  a. If you use 2.3 moles of Al, how many moles of  $\text{Al}_2\text{O}_3$  can you make? b. If you want 3.9 moles of  $\text{Al}_2\text{O}_3$ , how many moles of  $\text{O}_2$  are needed? 2.

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Chapter 12 Stoichiometry Packet Answer Key Chapter 12 Stoichiometry Packet Answers Chapter 12 Stoichiometry Packet Answer Key Answer:  $4.93 \times 10^{-5}$  L or 49.3  $\mu\text{L}$  In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol, and para -nitrophenol.

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Chapter 12 Stoichiometry 299. In the reaction represented by the equation  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ , how many grams of ... Answer the questions above, assuming we started with 30 grams of ammonium nitrate and 50 grams of sodium phosphate. Consider the following reaction:

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Chapter 12 Stoichiometry 12.1 The Arithmetic of Equations 12.2 Chemical Calculations 12.3 Limiting Reagent and Percent Yield. ... excess of 1800 is a logical answer. • The unit of the known (FSW 3 HP 2) cancels. • The answer has the correct unit (W). 3 Evaluate Does the result make sense?

Chapter 12

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13 Ch 12 Stoichiometry notes pkt. 1 mole =  $6.02 \times 10^{23}$  molecules (covalent) 1 mole =  $6.02 \times 10^{23}$  formula units (ionic) HOW MANY PARTICLES. 1 mole =  $6.02 \times 10^{23}$  atoms (monoatomic element) 1 mole = molar mass (grams) - HOW HEAVY. 1 mole = 22.4L for a gas at STP - HOW MUCH SPACE. Stoichiometry: Ex.  $\text{H}_2$  reacts with  $\text{N}_2$  to produce  $\text{NH}_3$ .

CHAPTER 11: STOICHIOMETRY

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