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Unit 6 - Test (Chemical Compound) In chemistry, a compound refers to an entity that consists of two or more atoms, with at least two of those atoms from different chemical elements. These atoms associate via chemical bonds.

Unit 6 - Test (Chemical Compound) - ProProfs Quiz

The lines/curves indicate the maximum amount of solute which can be dissolved in 100g of water at each temperature. We call this a saturated solution.. Any amount above the curve is said to be supersaturated; there is more solute than the solvent can hold, therefore the excess precipitates out/settles to the bottom.. Any amount below the line is called unsaturated, meaning more solute can be ...

Unit 6 | chemistry

UNIT 6: RADIOACTIVITY AND NUCLEAR CHEMISTRY 5. A skeleton of a mammoth was analysed and found to contain 21.2% of the carbon-14 found in living bone. The half-life of carbon-14 is 5730 years. What is the most likely year of the mammoth's death? Use  $\log(N/N_0) = -\lambda t$   $\lambda = \ln 2 / t_{1/2}$   $N/N_0 = 0.212$   $\ln(0.212) = -\lambda t$   $t = \ln(0.212) / (-\lambda) = \ln(0.212) / (-\ln 2 / 5730)$   $t = 5730 \times \ln(0.212) / (-\ln 2) = 10021.2$  years; hence find t This year  $\square$  t = year of death

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UNIT 6  $\square$  REDOX REACTIONS 6  $\square$  The oxidation number of an atom is the charge that would exist on an individual atom if the bonding were completely ionic  $\square$  In simple ions, the oxidation number of the atom is the charge on the ion: - Na<sup>+</sup>, K<sup>+</sup>, H<sup>+</sup> all have an oxidation number of +1 - Mg<sup>2+</sup>, Ca<sup>2+</sup>, Pb<sup>2+</sup> all have an oxidation number of +2 - Cl<sup>-</sup>, Br<sup>-</sup>, I<sup>-</sup> all have an oxidation number of -1

UNIT 6 REDOX REACTIONS - A-Level Chemistry

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GCSE Chemistry (Single Science) - BBC Bitesize

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