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The first step is to replace the molecules in the square brackets with the units of concentration. Using our previous example: Kc = [HI]2 / ([H2] [I2]) = (mol dm-3)2 / (mol dm-3)(mol dm-3) You can see that the numerator, (mol dm-3)2, is the same as the denominator, (mol dm-3)(mol dm-3) and so all the units cancel.

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Calculate Kc at this temperature. CH3COOH(l) + CH3CH2OH(l) CH3COOCH2CH3(l) + H2O(l) .1.00 a-. 00. ...v.s.s o. XY<S. +0. (4) (A2 BBoF 29) 1.90 mol of hydrogen and 1.90 mol of iodine were allowed to reach equilibrium at 710 K. The equilibrium mixture contained 3.00 mol of HI.

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3of 0.100 mol dm-3sodium butanoate. c) 1.00 g of potassium ethanoate is dissolved in 50 cm3of 0.200 mol dm-3ethanoic acid (K. a= 1.74 x 10. -5mol dm-3). 2) Calculate the pH of the following buffer solutions made by mixing an excess of weak acids with strong bases. a) 25 cm3of 0.5 mol dm-3methanoic acid (K.

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