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Thermodynamics Wyatt Tenhaeff Milo Koretsky

Department of Chemical Engineering Oregon

State University koretsm@engr.orst.edu 2 6.1

(a) The Clausius-Clapeyron equation: $dP_i \text{ sat}$

$P_i \text{ sat} \ln \frac{P_i \text{ sat}}{101 \text{ kPa}}$

$= -\frac{h_{i,vap}}{RT^2} \ln \frac{P_i \text{ sat}}{101 \text{ kPa}}$

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Department of Chemical Engineering Oregon
State University koretasm@enr.orst.edu 2 6.1

(a) The Clausius-Clapeyron equation: $dP/P = \frac{h_{vap}}{RT^2} \ln P$ sat
101 kPa $\ln \frac{P_2}{P_1} = \frac{h_{vap}}{R} \left(\frac{1}{T_1} - \frac{1}{T_2} \right)$
 $\ln \frac{101}{P_2} = \frac{h_{vap}}{R} \left(\frac{1}{T_1} - \frac{1}{T_2} \right)$
 $\ln \frac{101}{P_2} = \frac{h_{vap}}{R} \left(\frac{1}{T_1} - \frac{1}{T_2} \right)$
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Chapter 6 Solutions Engineering and Chemical Thermodynamics Wyatt Tenhaeff Milo Koretsky Department of Chemical Engineering Oregon State University koretsm@engr.orst.edu 2 6.1
(a) The Clausius-Clapeyron equation: $dP_i \text{ sat} = \frac{P_i \text{ sat}}{RT^2} \left(\frac{h_{i,v}}{R} - T \right) dT$ or $\ln \frac{P_i \text{ sat}}{101 \text{ kPa}} = \frac{h_{i,v}}{R} \left(\frac{1}{T} - \frac{1}{T_0} \right) - \ln \frac{P_i \text{ sat}}{P_i \text{ sat}(T_0)}$ so $P_i \text{ sat} = \dots$

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