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If NaCl is doped with 10 – 4 mol % of Sr Cl 2, the concentration of cation vacancies will be ($N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$) [CBSE AIPMT 2007] A $6.02 \times 10^{16} \text{ mol}^{-1}$

Solid State MCQs | Question – Paper 1

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Metallic solid conducts electricity in solid state but ionic solids do so only in molten state or in solution or metals conduct electricity through electrons and ionic substances through ions. Metallic solids are malleable and ductile while ionic solids are hard and brittle.

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Ans: It is a covalent or network solid. 1.8. Why are ionic solids conducting in the molten state and not in the solid state ? Ans: In the ionic solids, the electrical conductivity is due to the movement of the ions. Since the ionic mobility is negligible in the solid state, these are non-conducting in this state.

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